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CHE100 Stoichiometry Problems Instructor: S. Henzel $2 \text{ C}_3\text{H}_6\text{O}_2 + 7 \text{ O}_2 \rightarrow 6 \text{ CO}_2 + 6 \text{ H}_2\text{O}$. How many moles of CO_2 are produced when 8.652 moles of $\text{C}_3\text{H}_6\text{O}_2$ react? $8.652 \text{ mol C}_3\text{H}_6\text{O}_2 \times \frac{6 \text{ mol CO}_2}{2 \text{ mol C}_3\text{H}_6\text{O}_2} = 25.96 \text{ mol CO}_2$
22. How many moles of H_2O are produced when 50.75 g of O_2 react? $50.75 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{6 \text{ mol H}_2\text{O}}{7 \text{ mol O}_2} = 1.359 \text{ mol H}_2\text{O}$
23. How many moles of O_2 are requir...

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